

Cl isotope presentation

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Page 1: Diagrams illustrating the small isotopic effect of crystallization of halite.

Top: Ranges of ^{37}Cl expected for evaporation of seawater in a basin with continual replenishment of seawater, and in a basin with no replenishment of seawater after the initial input.

Bottom: Variation of ^{37}Cl in a batch of seawater evaporated until K and Mg salts appear. Note the small enrichment of halite (predicted as 0.24‰) relative to coexisting NaCl solution. The time scale is arbitrary. First appearance of halite and hexahydrate indicated by arrows.

Page 2: Diagrams illustrating the magnitude of isotopic fractionation resulting from linear diffusion of chloride ion at 25°C from an infinite source with $^{37}\text{Cl} = 0.0\text{‰}$ into pure water.

Top: Profiles of ^{37}Cl as functions of distance from source and time elapsed since initiation of diffusion.

Bottom: Profile of ^{37}Cl as a function of c_x/c_0 where c_x is the Cl concentration at distance x from the source, and c_0 is the source concentration. The curve is the same, regardless of time elapsed. Solutions with $^{37}\text{Cl} = -2\text{‰}$ have a Cl concentration that is 0.2 of the source concentration.

Page 3: Values of ^{37}Cl in near-surface sediments in southern Arizona, compared with the ^{37}Cl ranges of ideal marine evaporites.

Sediments from Safford Basin have positive values, some extreme for the Cl isotope system. If generated by diffusion, they represent residual chloride, probably the result of multiple cycles of wetting and drying with diffusive loss of chloride in each cycle.

Two negative ^{37}Cl values for soil from Tucson may indicate diffusion of chloride away from a near-0‰ source.





